I'm not a bot

Pb co3 2

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The license may not give you all of the permissions necessary for your intended use. For example, other rights such as publicity, privacy, or moral rights may limit how you use the material. Lead carbonate Names IUPAC name Lead(II) carbonate Other names Cerussite Identifiers CAS Number 598-63-0 Y ChemSpider 11234 ECHA InfoCard 100.009.041 EC Number 209-943-4 PubChem CID 11727 RTECS number OF9275000 UNII 43M0P24L2B Y CompTox Dashboard (EPA) DTXSID90883460 Properties Chemical formula PbCO3 Molar mass 267.21 g/mol Appearance White powder Density 6.582 g/cm3 Melting point 315 °C (599 °F; 588 K) (decomposes) Solubility in water 0.00011 g/(100 mL) (20 °C) Solubility product (Ksp) 1.46·10−13 Solubility insoluble in alcohol, ammonia; soluble in acid, alkali Magnetic susceptibility (χ) −61.2·10−6 cm3/mol Refractive index (nD) 1.804[1] Hazards GHS labelling: Pictograms Signal word Danger Hazard statements H302, H332, H360, H373, H410 Precautionary statements P201, P202, P260, P261, P264, P270, P271, P273, P281, P301+P312, P304+P312, P304+P340, P308+P313, P312, P314, P330, P391, P405, P501 Flash point Non-flammable Safety data sheet (SDS) External MSDS Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). N verify (what is YN ?) Infobox references Chemical compound Lead(II) carbonate is the chemical compound with the chemical formula PbCO3. It is a white, toxic solid.[2] It occurs naturally as the mineral cerussite.[3] Like all metal carbonates, lead(II) carbonate adopts a dense, highly crosslinked structure consisting of intact CO2−3 and metal cation sites. As verified by X-ray crystallography, the Pb(II) centers are seven-coordinate, being surrounded by multiple carbonate ligands. The carbonate centers are bonded bidentate to a single Pb and bridge to five other Pb sites.[4] Pb site in PbCO3, highlighting seven-coordination and the presence of one bidentate carbonate ligand for each Pb center. Lead carbonate is manufactured by passing carbon dioxide into a cold dilute solution of lead(II) acetate, or by shaking a suspension of a lead salt more soluble than the carbonate with ammonium carbonate at a low temperature to avoid formation of basic lead carbonate.[2] Pb(CH3COO)2 + [NH4]2CO3 → PbCO3 + 2 [NH4](CH3COO) Lead carbonate is used as a catalyst to polymerize formaldehyde to poly(oxymethylene). It improves the bonding of chloroprene to wire.[2] The supply and use of this compound is restricted in Europe.[5] A number of lead carbonates are known: White lead, a basic lead carbonate, 2PbCO3·Pb(OH)2 Shannonite, PbCO3·PbO Plumbonacrite, 3PbCO3·Pb(OH)2·PbO[6] PbCO3·2PbO Abellaite, NaPb2(OH)(CO3)2 Leadhillite, 2PbCO3·PbSO4·Pb(OH)2 ^ Pradyot Patnaik. Handbook of Inorganic Chemicals. McGraw-Hill, 2002, ISBN 0-07-049439-8 ^ a b c Carr, Dodd S. (2005). "Lead Compounds". Ullmann's Encyclopedia of Industrial Chemistry. Weinheim: Wiley-VCH. doi:10.1002/14356007.a15_249. ISBN 978-3-527-30673-2. ^ Inorganic Chemistry, Egon Wiberg, Arnold Frederick Holleman Elsevier 2001 ISBN 0-12-352651-5 ^ Sahl, Kurt (1974). "Verfeinerung der Kristallstruktur von Cerussit, PbCO3". Zeitschrift für Kristallographie. 139 (3–5): 215–222. Bibcode:1974ZK....139..215S. doi:10.1524/zkri.1974.139.3-5.215. ^ "EU law - EUR-Lex". ^ S.V. Krivovichev and P.C. Burns, "Crystal chemistry of basic lead carbonates. II. Crystal structure of synthetic 'plumbonacrite'." Mineralogical Magazine, 64(6), pp. 1069-1075, December 2000. "Archived copy" (PDF). Archived from the original (PDF) on 2009-05-21. Retrieved 2009-05-21.{{cite web}}: CS1 maint: archived copy as title (link) International Chemical Safety Card 0999 Retrieved from " Lead(II) carbonate is the chemical compound with the chemical formula PbCO3. 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Crystal structure of synthetic 'plumbonacrite'." Mineralogical Magazine, 64(6), pp. 1069-1075, December 2000. "Archived copy" (PDF). Archived from the original (PDF) on 2009-05-21. Retrieved 2009-05-21.{{cite web}}: CS1 maint: archived copy as title (link) International Chemical Safety Card 0999 Retrieved from " Please enable Javascript in order to use PubChem website.

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