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Class 10 science chapter 4 notes download
For example; Methane (CH4), Ethane (C2H4), Ethane (
number. Chlorine (—Cl) Chloro — 4. Students may access relevant study materials such as chapter notes, mind maps, diagrams, key questions, and a plethora of additional study resources by downloading the PDF links. The NCERT 10th class textbook's chapter study resources on carbon and its compounds can aid students in better test
preparation and in answering relevant guestions. H2SO4. Types of Covalent Bond: When a single pair of electrons are shared between two atoms in a molecule. The revision notes help you revise the whole chapter in minutes. This result in formation of a radial structure called micelles. Saponification Reaction: Reaction of esters
 with sodium hydroxide, gives alcohol and sodium salt of carboxylic acid (soap). In this chapter, the students learn about carbon and its covalent bond along with the IUPAC process based on the associated functional groups. C(s) + O2(g) 
air or oxygen. This type of bond is called a covalent bond. This type of bond is also formed between other compounds, such as in H2, O2, N2, etc. Allotropes of Carbon atom is bonded to another four carbon atom is bonded to another four carbon atom is bonded to another four carbon atom.
to three other carbon atoms in the same plane, giving rise to a hexagonal array. C-60 Buckminsterfullerene; In C-60 Buckminsterfullerene; arbon atoms are arranged in the shape of a football. Versatility Nature of Carbon Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football. Versatility Nature of Carbon atoms are arranged in the shape of a football.
can bond with a single bond, a double bond, or a triple bond. Carbon forms strong bonds with elements, and because of its small size, the nucleus is able to hold upon the shared pair of electrons. In this, the two carbon atoms are connected by double bond. Nomenclature of Carbon Compounds (i) Identify the number of carbon atoms in compounds.
Due to this, a lot of soap gets wasted and cleansing action gets reduced to a larger extent. Chemical Properties Reaction with sodium R
Reactions of ethanoic acid: Saponification: Esters in the presence of acid or base react to give back alcohol and carboxylic acid is called saponification. Carbon shows three allotropic forms: Note: Diamond is the hardest substance whereas graphite is very soft. This gives an opportunity to other tails to stick to oil. In order to attain the noble gas
configuration, carbon should either gain 4 electrons or lose 4 electrons or can share it's 4 electrons with some other element. Halogens, in IUPAC, are written as Prefixes, e.g., Compound With two carbons and one chloro group is named as chloroethane (CH3CH2CI). Detergents: Chemically, detergents are sodium salts of sulphonic acids, i.e.,
detergents contain a sulphonic acid group (—S03H), instead^of a carboxylic acid group (—COOH), on one end of the hydrocarbon. Example: Isomers of butane (C4H10) Homologous Series: Series of organic compounds having the same functional group and chemical properties and successive members differ by a CH2 unit or 14 mass units are known
as Homologous series. A long hydrocarbon part which is hydrophobic (water repelling) in nature and a short ionic part which is hydrophobic (water repelling) in nature. No. Functional Group Prefix Suffix 1. All the members of a series show similar chemical properties. Three carbon atoms (3-C) Prop + ane 4. Triple Covalent Bond: When three pairs of
electrons are shared between two atoms in a molecule. The physical properties vary among the members of homologous series due to difference in their molecular mass. Carbon has Three Main Allotropes Diamond: In this, carbon, an atom is bonded to four other atoms of carbon forming three-dimensional structures. (d) Substitution Reaction (a)
Combustion Reaction: A chemical reaction in which a substance burns in the presence of air or oxygen is called combustion reaction. Hardness of water is due to the presence of bicarbonates, chlorides and sulphate salt of calcium and magnesium. 12. Pure acetic acid is called glacial acetic acid SODIUM NaSODIUM ETHANOATE AND
HYDROGENSODIUM CARBONATEREACTS WITHPRODUCTS1.2.SODIUM ETHANOATE, CARBONDIOXIDE AND WATER3.SODIUM BICARBONATESODIUM BIC
Ketone (-CO-) — one 8. It is simpler to learn things through written notes. The recurring questions and the ones that need to be addressed first are covered by the critical questions. Ethene - C2H4 Propene - C3H6 Homologous series: A family of organic compounds having the same functional group, similar
chemical properties and the successive (adjacent) members differ by a CH2 unit or 14 mass unit. → Alkanes are saturated hydrocarbons. An unsaturated hydrocarbon with a double bond having two carbon atoms is named as Ethene. Cleansing action of soap: The dirt is generally held to the surface of a dirty cloth by a thin film of oil or grease. The
functional group present in it is carboxylic acid - COOH. - It is series of compounds in which the some functional group substitutes for the hydrogen in a carbon chain. Difference between soaps and detergents (i) These are sodium or potassium salts of long chain fatty acids. - Carbon and its Compounds Support MaterialThere are
a ton of study resources available everywhere, so it will take time and effort to narrow down the selection of the best study materials for carbon atoms. These acids are present in the form of their esters along with glycerol (alcohol containing three
hydroxyl groups). Download revision notes for Carbon and its Compounds class 10 Notes and score high in exams. The PDFs include all the key ideas and subjects and are carefully selected with the updated CBSE syllabus in mind. The topics are thoroughly covered and presented understandably in the metal and non-metals PDF. With only a few clicks
on the links and OTP authentication, educators and students may access the PDFs. Once the information has been downloaded to the device, you may access the PDFs whenever and wherever you like. Ethanoic acid reacts with alcohols in the presence of cone. The simplest molecule formed by sharing of electrons (i.e., covalent bonds), can be
represented by electron dot structure. If a suffix is added, then final 'e' is removed from the name e.g. methanol (methane-e = methan + ol). Chemical Properties of Carbon Compounds 1. (iii) In the case of a functional group present, write the prefix or suffix of the functional group as given below. The answer to all these questions lies in carbon and its
compounds.NCERT Class 10 Chapter 4 Notes Carbon CompoundsChemical Properties of Carbon Chemical 
Solutions Subject WiseNCERT Class 10 Science Chapter 4 Notes Carbon and its compounds are provided to revise all the important concepts given in this chapter. Ethanol is one of the main components of alcoholic drinks.
When water is agitated, the oily dirt tends to lift off from the dirty surface and dissociates into fragments. (ii) Ionic part of detergent is —COO-Na+ (ii) Ionic part of the soap is —COO-Na+ (ii) Ionic part of detergent is —COO-Na+ (ii) Ionic part of detergent is —COO-Na+ (ii) Ionic part of detergent is —COO-Na+ (iii) Ionic 
are called hydrocarbon. Create NowQuestion Bank, Mock Tests, Exam Papers, Notes Install Now CBSE Class 10 Science Notes for Quick Revision. The ionic-end of soap called head interacts with water and the carbon chain
called tail interacts with oil. In hard water, soaps form scum with water. It has a very low melting point (156 K) and low boiling point (351 K or 78°C). Saturated hydrocarbon is 6. Example, Saturated hydrocarbon having one carbon atom is named as
Methane. H2SO4 acid absorbs water molecules from the alcohol molecules and acts as a strong dehydrating agent. This increases the number of carbon dioxide and water, and release heat and light energy. Students can utilize these notes to revise major
concepts while preparing for the exams. Physical Properties It is colourless, inflammable liquid. The rules for nomenclature are as follows: (i) Identify the number of electrons contributed by each atom for sharing. Physical Properties of Ethanoic acid • Colourless liquid having sour
 taste and have smell of vinegar. Thus, to systematize the nomenclature of organic compounds IUPAC (International Union of Pure and Applied Chemistry) has given certain rule which is as follows: 1. Uses: As solvent, as antiseptic (tincture iodine), as anti-freeze in automobiles. H2SO4 as catalyst gives a sweet-smelling substance known as Esters,
called Esterification reaction. All the members of a given homologous series have the same functional group. Unsaturated hydrocarbons burn with a vellow smoky flame, • They have same chemical properties but show gradual change in physical properties. Diamond is used for grinding and polishing of 'hard materials and graphite is used as a
lubricant. Each element produces its own characteristic colour. Formation of coal and petroleum coal has been formed from remains of trees and plants which have been subjected to various biological and geological processes for millions of years ago, wherein they were crushed and buried down the earth layer by layer by earthquake or
 volcanoes. Petroleum has been formed from remains of plants and animals in the sea which have been attacked by bacteria under high pressures and converted into oil and gas, which are trapped between the rocks. Oxidation When oxygen is added to the compound, the reaction is called oxidation reaction. Here we have given NCERT Class 10 Science
 Notes Chapter 4 Carbon and its Compounds. Versatile nature of Carbon, Hydrocarbons, Isomerism, Homologous series, Functional groups, Nomenclature of functional groups, Nomenclature of functional groups, Nomenclature of functional groups. Ethanol: (or alcohol) Colourless liquid, soluble in water, and has a distinct smell and burning taste. Saturated Compounds of carbon which have only single bonds
 between the carbon atoms are called saturated compounds e.g., Ethane, Propane, Butane etc. Topics in the Chapter • Introduction • The Covalent Bond • Versatile nature of Carbon Compounds • Soaps and Detergents Introduction → Compounds are of two types: (i) Organic
Compounds (ii) Inorganic Compounds are made up of Carbons and form the basis of all living organisms. The bond formed by mutual sharing of electron pairs between two atoms in a molecule is known as Covalent Bond. Recall that electro-negativity decreases as atom moves further away from fluorine on the periodic chart. Six
carbon atoms (6-C) Hex + ane 2. Two carbon atoms (2-C) Eth + ane 3. Saturated hydrocarbon having two carbon atoms is named as Ethane. It is also called hydrogenation of vegetable oils. Pure ethanoic acid (Esterification reaction) \mathrm{CH} 3 \mathrm{COOH}+\mathrm{CH} 3
\mathrm{CH} 2 \mathrm{OH} \rightarrow \mathrm{COOCH} 2 \mathrm{COOCH} 3 \mathrm{COOCH} 2 \mathrm{COOCH} 2 \mathrm{COOCH} 3 \m
carbonates)$\begin{aligned} & 2 \mathrm{COOH}+\mathrm{CO}} 3 \rightarrow \mathrm{COOH}+\mathrm{CO}} 3 \mathrm{COONa}+\mathrm{CO}} 3 \mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{COONa}+\mathrm{C
\mathrm{O}+\mathrm{CO} 2\end{aligned}$Soaps and DetergentsSoaps are the long-chain carboxylic acids of sodium or potassium. 2. (i) Ionic (hydrophobic) part • Most dirt is oily in nature and hydrophobic end attaches itself with dirt and the ionic end is surrounded with molecule of water. Carbon can
bond with four carbon atoms, monovalent atoms, oxygen, nitrogen and sulphur. • Boiling point is 391 K. But it requires a large amount of energy to remove four electrons. Or why do plants need carbon dioxide to survive? Example : Oxygen molecule, Carbon Dioxide molecule. Note : Cl is named as prefix Chloro; Br as Bromo; NH2 as Amino and NO2 as
Nitro. Bromine (-Br) Bromo - 5. Transportant Questions (View) Below, we have provided Class 10 Science Mind maps of the related concepts in Carbon and its Compounds. Soaps and Synthetic Detergents: Soaps and Synthetic Detergents: Soaps and detergents are substances used for cleaning. (ii) Reaction
branched structures, carbon compounds are present in cyclic form. Identify the number of carbon atoms in the compound. 2. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix. Functional group is indicated either by prefix or suffix or suffix or suffix and the prefix or suffix or suffi
successive members of a homologous series differ by 14 u. Process of converting vegetable oil into solid fat (vegetable quee) is called Hydrogenation of Oil. Students must go through each topic in the carbon and its compounds class 10 notes as easily and effectively as possible, with the help of the NCERT Notes for Class 10. Also, students can refer
toNCERT Class 10 Chapter 4 Notes Carbon and its Compounds Bonding in Carbon: The Covalent BondWe know that an ionic bond is formed when an atom loses an electron and another atom gains and electron and another atom gains atom gains and electron and another atom gains at electron and atom gains at electron and atom gains at electron and atom gains at electron atom gains at electron atom gains at electron and atom gains at electron at
requirements or even gain 4 electrons due to repulsion between the electrons. Therefore, carbon forms bonds by sharing its 4 electrons with another carbon atom, hence completing its octet. Identify the functional group S. Have you ever wondered why fossil fuels are used to power vehicles? 5. For example, isomers of butane. Oxidizing Agent: Some
 substances are capable of adding oxygen to others, are known as Oxidising Agent. 13. It is dark solid at room temperature. If you have any query regarding NCERT Class 10 Science Notes Chapter 4 Carbon and its Compounds, drop a comment below and we will get back to you at the earliest. For example; F2, Cl2, H2 etc. Oxidation: Oxidation of
 ethanol in presence of oxidizing agents gives ethanoic acid. Carbon forms bond with many elements, such as oxygen, hydrogen, nitrogen, sulphur, chlorine, etc, thus forming compounds which involve single bonds are called saturated carbon compounds. Unsaturated Carbon
Compounds Carbon compounds which involve double or triple bonds are called unsaturated Carbon Compounds. Naming of Saturated Carbon Compounds Hetero AtomClass of
CompoundsFormulaCl/BrHalo-(chloro, bromo)alkane-Cl, -BrOxygenAlcohol-OHAldehyde-CHOKetone-COCarboxylic acid-COOHNomenclature of Carbon atoms in the compound and based on the number of carbon atoms in the compound and based on the number of carbon atoms in the compound and based on the number of carbon atoms in the compound and based on the number of carbon atoms.
 with the electrons of other carbon atom or with other element and attain noble gas configuration. Its common name is ethyl alcohol. Scum: The insoluble precipitates formed by soap molecule when they react with calcium and magnesium ions present in hard water. The negatively charged heads present in water prevent the small globules from
coming together and form aggregates. The element that replaces hydrogen is called a hetero atom. Alkynes: General formula is CnH2n-2, where n = number of carbon atoms. In this, the two carbon atoms are connected by triple bond. In addition to combustion, oxidation can also be: brought about by some substances which are capable of giving
oxygen to others, i.e., Oxidising agents, e.g., Acidified K2Cr207 (Potassium dichromate) and alkaline KMn04 (Potassium permanganate). Homologous series of Alkanes, Alkenes and Alkynes Characteristic of Homologous Series The successive members in homologous series of Alkanes, Alkenes and Alkynes Characteristic of Homologous Series The successive members in homologous series of Alkanes, Alkenes and Alkynes Characteristic of Homologous Series The successive members in homologous Series The successive members in homologous series of Alkanes, Alkenes and Alkynes Characteristic of Homologous Series The successive members in homologous Series The successive members and the successive members are successive members are successive members and the successive members are successive members and the successive members are successiv
only a single bond. Soaps are effective with soft water only and ineffective with hard water. Substitution Reaction: (b) Oxidation Reaction: The addition of oxygen in a compound upon combustion is called
oxidation. (iv) Soaps are biodegradable. Electron Dot Structure of Unsaturated Hydrocarbons Names, molecular formulae and structure formulae and structure formulae of unsaturated hydrocarbons (Alkenes and Alkynes): Carbon
Compounds on the Basis of Structure (i) Straight (unbranched) chain - These three above compounds has same molecular formula but different structural isomerism. Fullerenes are called structural isomerism or a soccer ball like. Soap: Sodium or
 potassium salts of long chain fatty acids is called Soap. CH3CH2OH(l) + O2(g) \rightarrow CO2(g) + H2O(l) + Heat and light Carbon burns in air or oxygen to give carbon dioxide and heat and light. Hydrocarbons: a Large number of hydrocarbons can be classified as: Note: In an open chain, the name of parent chain is derived from the root word and suffix ane
ene or yne is added depending on the type of bond present in a chain: Important: No alkene or alkyne is possible with single carbon atom because double or triple bond is not possible between carbon atom because double or triple bond is not possible between carbon atom.
in our surroundings. Physical Properties It is a colourless, pungent-smelling liquid. Loss of 4 electrons (to attain duplet, i.e., 2 electrons like He atom in C4+ cation) is difficult as it requires a large amount of energy to remove four electrons. For example, a 4-carbon chain with a ketone group would be named Butane - 'e' = Butan + 'one' =
Butanone.Chemical Properties of Carbon CompoundsCombustionCarbon reacts with oxygen, giving out carbon dioxide, heat and light. \\&\mathrm{O} 2 \rightarrow \mathrm{CO} 2+\mathrm{C} +\mathrm{CO} 2+\mathrm{C} +\mathrm{C} +\m
light } \\ & \mathrm{CH} 3 \mathrm{CH} 2 \mathrm{O} + \text{ heat and light } \end{aligned}$Why substances burn with a flame because a flame is produced when gaseous substances burn with a flame?Substances burn with a flame because a flame is produced when gaseous substances burn and glow. 2C2H5OH + 2Na \rightarrow 2C2H5ONa (Sodium flame)
ethoxide) + H2 (g) Ethanol on dehydration in the presence of come. Nomenclature of Organic Compounds: It is difficult to remember millions of compounds by their individual common name. Soap or detergent is attracted both by the greasy dirt and water. Versatile
 Nature of Carbon The two characteristic properties of carbon element which lead to the formation of large number of compounds: • Catenation: Carbon can link with carbon atoms by means of covalent bonds to form long chains, branched chains and closed ring. → In hydrocarbon chain, one or more hydrogen atom is replaced by other atoms in
 accordance with their valancies. (i) Catenation: The self linking property of an element mainly carbon atom through covalent bonds to form stable multiple bonds (double or triple) with itself and with the atoms of other elements. It is a good conductor of
electricity and used as a lubricant. Burning of coal and petroleum emits oxides of sulphur and nitrogen which are responsible for acid rain. 2. OXIDATION: Alcohols can be converted to carboxylic acids by oxidizing them using alkaline potassium permanganate or acidified potassium di-chromate (they add oxygen to the reactant, thus are called
oxidizing agents). Example :Nitrogen,C2H2 Properties of Covalent Bonds.- Covalent Bonds.- Covalent Bond of Carbon always forms covalent bond. Unsaturated hydrocarbon add hydrogen in the presence of catalyst palladium or nickel. These are
short ideas on carbon and its compounds, with an explanation of each topic and formula. • Volatile liquid with low boiling point of 351 K. Electron Dot Structure: The electron dot structures provides a picture of bonding in molecules in terms of the shared pairs of electrons and octet rule. Their charged ends do not form scum and thus are used for
 cleansing purposes. Class 10 Science: Chapter Wise CBSE Notes 10 NCERT Solutions Subject WiseNCERT Solutions for Class 10 MathsNCERT Solutions for Class 10 MathsNCERT Solutions for Class 10 NCERT Solutions for Class 10 MathsNCERT Solutions for Class 10 Exemplar Solutions for Class 10 MathsNCERT Solutions for Cl
Compounds...Download the app to get CBSE Sample Papers 2024-25, NCERT Solutions (Revised), Most Important Questions, Previous Year Question Bank, Mock Tests, and Detailed Notes. Nomenclature of carbon compounds. Five carbon
 atoms (5-C) Pent + ane 6. This alcohol is called denatured alcohol. Ethanoic Acid is colloquially called acetic acid, its 5-8% solution is called vinegar and is used widely for cooking purposes. These hetero atoms and the group containing them impart chemical properties to the compound and hence are called functional groups. Homologous and the group containing them impart chemical properties to the compound and hence are called functional groups. Homologous and the group containing them impart chemical properties to the compound and hence are called functional groups.
Series It is a series of compounds in which the same functional group substitutes for hydrogen in a carbon chain. We hope the given CBSE Class 10 Science Notes Chapter 4 Carbon and its Compounds Pdf free download will help you. Alcohols can be converted to carboxylic acid in presence of oxidizing agent alkaline KMnO4 (potassium permangnate)
or acidic potassium dichromate. Miscible with water in all proportions. Scum: The magnesium and calcium salts present in hard water reacts with soap molecule to form insoluble products called scum, thus obstructing the electrons so that each atom contributes one electron to a single bond between each atom. The
topics covered in these notes are: definitions, bonding in carbon—the covalent bond, versatile nature of carbon compounds, chemical properties of carbon compounds, some important carbon compounds—ethanoic acid, soaps, and detergents.
So, its configuration is K-2, L-4. The chapter has enough weight on the test paper, thus it's important to prepare for it using suitable resources. It could lose four electrons forming C4+ cations. Note: Compounds of carbon with double bonds and
triple bonds are called as unsaturated compounds while those with carbon-carbon single bonds are called acetic acid (CH3COOH) is a colourless liquid. Example: Alkaline KMnO4 (or KMnO4—KOH) Acidified K2Cr2O7—H2SO4) KMnO4 -
 Potassium permanganate K2Cr2O7 - Potassium dichromate 3. The atmosphere has 0.03% of carbon dioxide. Note: (c) Addition reaction, e.g., Hydrogenation of vegetable oils as shown in the reaction below: (d)
 Substitution Reaction: The reactions which involve the replacement of an atom or group of atoms from a molecule by another atom without any change in structure in the remaining part of the molecule. Redraw the dots so that electrons on any given atom are in pairs wherever possible. Electron dot structure of a saturated carbon compound, ethane is
as follows: Electron dot structure of an unsaturated carbon compound, ethene is as follows: TRY DRAWING THE ELECTRON DOT STRUCTURE OF ETHYNEFormulae and Structures of Saturated Compounds having identical
molecular formula but different structures. During the revising sessions, the notes are helpful. The mindmaps help in connecting ideas and showing how related topics are to one another. Vegetable glee Vegetable fats which are harmful for health. Addition
 Reaction: Addition of dihydrogen with unsaturated hydrocarbon in the presence of catalysts such as nickel or platinum or palladium are known as Hydrogenation (addition) reaction. Found in .02% in form of minerals an 03% in form of minerals are carbon based. Covalent Bond in Carbon in the presence of catalysts such as nickel or platinum or palladium are known as Hydrogenation (addition) reaction.
configuration is 2, 4. The best app for CBSE students now provides Carbon and its Compounds class 10 Notes latest chapter wise notes for quick preparation of CBSE board exams and school based annual examinations. Chemical Properties of Ethanol 2Na + CH3CH2ONa+ + H2 This reaction is used as a test for ethanol by evolution
 of H2 gas (Burn with pop sound). 1. Unsaturated Compounds of carbon which contain one or more double or triple bonds between carbon atoms are called unsaturated compounds e.g., Ethene, Propene, Butyne, etc. Structural Isomerism: Compounds having the same molecular formula but different structures are called Structural isomers.
Three allotropic forms of carbon: Diamond Graphite Fullerenes Hydrocarbons. Organic compounds of carbon and hydrocarbons. Ethanoic acid reacts with bases to form its salt and water. Double bond (=) — ene 2. It is a neutral substance, so it does not have any effect on either blue litmus solution or red litmus solution. Repeat
steps 4 and 5 as needed until all octets are full. It is only between two carbon atoms. These electrons contributed by the atoms for mutual sharing in order to acquire the stable noble gas configuration is called covalency of that atom. It is miscible with water in all proportions. Sample Paper all are made available through the best app for CBSE
students and myCBSEguide website. Create question paper PDF and online tests with your own name & logo in minutes. - There are two types of Hydrocarbons. The cloth needs to be mechanically agitated to remove the dirt particles from the cloth. Note: Combustion is always an EXOTHERMIC reaction, e.g., Remember: Saturated hydrocarbons
generally give clean flame whereas unsaturated hydrocarbons give sooty flame (because carbon content in these, and hence carbon shows incomplete combustion and appears as soot). • Soap micelles helps to dissolve dirt and grease in water and cloth gets cleaned. This combustion is an oxidation process. - These
molecules are generally poor conductor of electricity since no charged particles are formed. Two important property of carbon to form bond with other atoms of carbon. Ethanoic Acid: CH3COOH Common Name: Acetic Acid. Its consumption in
 small quantities causes drunkenness and can be lethal. Most carbon compounds are poor conductors of electricity, due to the absence of free electrons and free ions. Ethanol is highly soluble in water. How are plastics made? This scum create difficulty in cleansing action. Compound Carbon atoms may be linked by single, double or triple bonds
Allotropes of carbon: The phenomenon by means of which an element can exist in two or more forms, with similar chemical properties but different physical properties are called allotropy and the different physical properties but different physical properties are called allotropy and the different physical properties but different physical properties are called allotropy and the different physical physical properties are called allotropy and the different physical phys
hydrocarbons Methane - CH4 Ethane - C2H6 Alkenes General formula - CnH2n Unsaturated hydrocarbon. For instance, the ALCOHOLS: OH, , , The successive member differs by one -CH2 unit and 14 units of mass. Therefore, it should either lose or gain 4 electrons to achieve the noble gas configuration and become stable. Chemical properties of
carbon compounds: Main properties of carbon compounds are: (a) Combustion Reaction (b) Oxidation Reaction (c) Addition Reaction type is covered by the worksheets and DoE worksheets, and answering these questions following the question
study schedule will improve your comprehension of the chapter. Teachers and students can prepare the chapter extensively with the help of the questions and other practice resources provided. -comparatively weaker intermolecular forces, unlike ionic compounds. Its electronic configuration is 2, 4. The tails stick inwards and the heads outwards.
 Allotropic forms of Carbon: The phenomenon of existence of an element in two or more forms which have different physical properties but identical chemical properties but identical chemical property of self-linkage of carbon
atoms by means of covalent bonds to form straight chains, or the rings of different sizes (as shown below): (b) Tetracovalency: Due to small size, and presence of four valence electrons, carbon can form strong bonds with other carbon atoms, hydrogen, or sulphur, etc. → The bond formed by sharing of electron
pair between two atoms are known as covalent atoms. It is difficult to hold four extra electron and would require large amount of energy to remove four electrons. The great strength of the carbon-carbon bond. Names which are given on this basis are popularly known as IUPAC name. In spite of its small amount available in nature, carbon is a
 versatile element as it forms the basis for all living organisms and many things which we use. And 100% pure acetic acid (C17H33COOH), Palmitic acid (C17H35COOH), Palmitic acid (C17H31COOH), etc. It is also used
for making jewellery. H2SO4 acid to form sweet smelling compounds called esters. Important: Replacement of hydrogen atom by a functional group is always in such a manner that valency of carbon remains satisfied. It cannot lose 4 electrons to form C4+ ion having Helium gas (2) configuration because this cation would be highly unstable due to a
large amount of energy required to remove four electrons from the carbon atom. Saturated fatty acids are harmful for health and oils with unsaturated fatty acids are harmful for health and lethanol consumption of dilute
ethanol causes serious health issues and intake of pure alcohol is lethal. CHEMICAL PROPERTIES OF ETHANOL reacts with Sodium to form an ionic bond It could gain four electrons forming C4- cation. But the members have different
physical properties. 2.TETRA VALENCY: Having a valency of 4, carbon atom is capable of bonding with atoms of oxygen, hydrogen, nitrogen, sulphur, chlorine and other elements. Triple bond (=) — yne 3. It reacts with carbonate and hydrogen, nitrogen, sulphur, chlorine and other elements.
release CO2 gas. The revision notes covers all important formulas and concepts given in the chapter. Cleansing Action of Soaps and Detergents: Both soaps and detergents cantains two parts. The NCERT books comprehensively study all the important concepts and topics, but students require extra study material for exam preparations. (iv)
Detergents are non-biodegradable. These esters, called 'glycerides' are present in fats and oils of animal and vegetable origin. Sooty flame is seen when unsaturated hydrocarbons are burnt. Example: C17H35COONa+ • Soaps are effective only in soft water. Saturated hydrocarbons are burnt. Example: C17H35COONa+ • Soaps are effective only in soft water.
 yellow sooty flame in limited supply of air. The gas and kerosene stove used at home has inlet for air so that, burnt to given clean blue flame. It is a good organic solvent. Carbon attains the noble gas configuration by sharing its valence electrons with other atoms. Such mutual sharing of electrons between atoms to attain a stable noble gas
 configuration is called Covalent bonding. Covalent bonding. Covalent bonding. Covalent bond formed between two atoms by sharing of valence electrons between two atoms so that each atom acquires the stable electronic configuration of the nearest noble gas. All the members of homologous series shows similar chemical properties. Nomenclature of Organic
Compounds Trivial or common names: These names were given after the source from which the organic compounds are used as fuels because they
burn in air releasing lot of heat energy. 6. One carbon atoms (1-C) Meth + ane 2. The chemical properties are imparted by the functional group thus all members have similar chemical properties. IUPAC name: International Union of Pure and Applied Chemistry gave following rules for naming various compounds:
and write the word root corresponding to it. Functional Group: An atom or group of atoms present in a molecule which largely determines its chemical properties are called Functional Group. 11. • Detergents are effective in both hard and soft water.
higher than saturated hydrocarbon which does not get completely oxidized in air. The smaller size of carbon atoms enables nucleus to hold the shared pair of electrons strongly, thus carbon atoms. Hydrocarbons are further
classified as Saturated and Unsaturated Carbon Compounds are hydrocarbons in which hydrogen and carbon have single bonds. • Noble gas configuration of Carbon - Carbon is tetravalent, it does not form ionic bond by either losing four electrons (C4+) or by gaining four electrons (C4+). Click the links to
get the carbon and its compounds class 10 PDF after registering as a user. These are the Carbon and its Compounds by replacing one hydrogen atom or atoms, e.g., -Cl.
 Functional group, replacing the hydrogen is also called as heteroatom because it is different from carbon, and can be nitrogen, sulphur, or halogen, etc. Since it requires four electrons, carbon is said to be tetravalent. e.g., a ketone with three carbon atoms is named as: Propane - e = Propan + 'one' = Propanone. It's melting point is 290 K and the
boiling point is 391 K. Ethanoic Acid (CH3COOH): Commonly known as Acetic acid. Due to presence of small amount of nitrogen and sulphur, coal and petroleum produces carbon dioxide with oxides of nitrogen and sulphur which are major pollutant. In cleansing, the hydrocarbon tail attaches itself to oily dirt. Decide how many valence (outer shell)
electrons are possessed by each atom in the molecule. It requires, 4 electrons to achieve the inert gas electronic configuration. (ii) Functional group is indicated by suffix or prefix. Example: CH3—(CH2)11—C6H4—SO3Na. Hard and Soft Water: Water that does not produce lather with soap readily is called Hard water and which produces lather with
soap is called Soft Water. SULPHURIC ACID)ESTER AND WATER5.NAOHSODIUM ETHANOATE AND WATERethonol of ester - ethnoic acid as catalyst and form sweet smelling compounds called ester. Hydrolysis On heating with an acid or a
base the ester forms back the original alcohol and carboxylic acid. 4. It is capable of bonding with four other atoms of carbon or some other heteroatoms with single covalent bond as well as double or triple bond. They are effective with both soft as well as hard water. An ionic part (hydrophilic) and a long hydrocarbon chain (hydrophobic) part
constitutes the soap molecule. Long chain of carboxylic acid, hydrophobic ionic end, hydrophobic ionic end attaches itself with dirt, while the ionic end is surrounded with molecules of water. Chemical Properties (i) Esterification Reaction: Reaction of Soaps Most dirt is oily in nature and the hydrophobic ionic end, hydrophobic ionic end, hydrophobic ionic end attaches itself with dirt, while the ionic end is surrounded with molecules of water. Chemical Properties (i) Esterification Reaction: Reaction of Soaps Most dirt is oily in nature and the hydrophobic ionic end, hydrophobic ionic end, hydrophobic ionic end, hydrophobic ionic end is surrounded with molecules.
 ethanoic acid with an alcohol in the presence of a few drops of conc. However, it is difficult for carbon to gain or lose four electrons because this anion would be highly unstable due to a large amount of energy required to overcome the
forces of repulsion between the four electrons being added and the six electrons already present in the carbon atom. This was done to maintain uniformity throughout the world. Straight Chain Branched Chain Rings (ii) Tetravalent Nature: Carbon has valency of four. Thus, in hard water, soap does not give lather while detergent does. • By use of
detergent, insoluble scum is not formed with hard water and cloths get cleaned effectively. • Saturated hydrocarbon generally burn in air with blue and non-sooty flame. The cleansing action of detergent is considered to be more effective than a soap. ALKANE: Unsaturated Compounds: hydrocarbons in which hydrogen and carbon have double or
triple bonds. 9. The derivations in the NCERT Notes Class 10 Science chapter 4. No Number of Carbon Atoms Word Root (-) (Suffix) Single bond 1. Used as raw material for the preparation of acetyl chloride and esters. For example; N2 etc. But it would be difficult for the nucleus with six protons to hold on to
ten electrons. Used as vinegar. Thus, the oily dirt is removed from the object. Alkane - CnH2n+1CHO Carboxylic acid - CnH2n+1CHO Car
homologous series differs by a CH2 unit. Graphite: In this, each carbon atom is bonded to three other carbon atoms. Then write the name of the parent compound: Chemical properties of Ethanol (C2H5OH) compound is a colourless liquid at room temperature. These notes can be helpful for CBSE board exam preparation. These notes give you
a basic idea of the key features of carbon and its compounds. H2SO4 acid at 443 K forms ethene gas. Bonding in carbon = 6 An electrons in L shell. Note that in some structures there will be open octets (example: the B of BF3), or atoms which have ten electrons in L shell.
 (example: the S of SF5) 6. 	o Alkenes and Alkynes are unsaturated hydrocarbons. With the help of worksheets, important questions, learning activities, and additional study materials students can get an in-depth understanding of carbon and its compounds class 10. Below, we have provided the links to downloadable PDFs of DoE Worksheets for class
10 Science to practice more questions. Esters are used in making perfumes and flavouring agents. NOMENCLATURE: Nomenclature is the system of assignment of names given to organic compounds. Heteroatom and Functional Group In hydrocarbon chain, one or more hydrogen atoms can be replaced by other atoms in accordance with their
 valencies. Diamond has a three-dimensional rigid structure but graphite has a hexagonal sheet layer structure. Example: Alcohols - CH3OH, C2H5OH, C3H7OH, C3H7
save your time during stressful exam days. Revisions Notes for class 10 Science Download as PDFCBSE Class 10 Science Revision Notes Chapter 4 Carbon and its compounds Carbon is a versatile element. Cleansing Action of Soaps and Detergents: The cleansing action of soaps and detergents follows the same principle. NaOH + CH3COOH -
CH3COONa + H2O (iii) Reaction with carbonates and hydrogen carbonates : 2CH3COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CHH2COOH + NaHCOH2 → CH3COONa + H2O + CO2 CH3COONa + H2O + CO2
of carbon atoms. The members of a homologous series show a gradation in physical properties. Hence, carbon shows TETRACOVALENCY. (c) Multiple Bonds or triple bonds, (i.e., double bonds or triple bonds) with other elements as well as with its own atoms. Preparation of Soap: When an oil or
fat (glyceride) is treated with sodium hydroxide solution, it gets converted to sodium salt of the acid (soap) and glycerol. For example; Cone. Detergents are ammonium or sulphonate salts of carboxylic acids with long chain. 8. It is the hardest substance and an insulator. Physical Properties
of Covalent Compounds - Covalent compounds have low melting and boiling points as they have weak intermolecular force. This reaction is known as Saponification Reaction. The melting point of pure ethanoic acid is 290 K and hence, it often freezes in cold climate so named as glacial acetic acid. 15. The hydrocarbon part of the soap molecule links
itself to the oily (dirt) drop and ionic end orients itself towards water and forms a spherical structure called micelles. Versatile Nature of carbon, Catenation Tetravalent nature. According to new CBSE Exam Pattern, MCQ Questions for Class 10
Science pdf Carries 20 Marks. hydrocarbons with double covalent bonds are alkenes and those with triple covalent bonds are alkynes (with —C = C—) are saturated compounds. It is done in the presence of an oxidising agent.$\mathrm{CH} 3
\mathrm{CH} 2 \mathrm{OH} \xrightarrow[\text { OrAcidified } 2 \mathrm{COH}$Addition ReactionWhen hydrogen is added to an unsaturated compound in presence of a catalyst such as palladium or nickel, giving rise to saturated hydrocarbons, the
reaction is called addition reaction. Substitution Reaction. Some Important Carbon Compounds - Ethanol and Ethanol is commonly called alcohol, it is a good solvent which is soluble in water also and is used as medicine to
treat cough.Reactions of Ethanol:(Reaction with sodium)2Na + 2CH3CH2OH $\rightarrow$ 2CH3CH2O-Na+ + H2(Reaction to give unsaturated hydrocarbon)$\mathrm{CH} 2 \mathrm{CH} 2 \mathrm{CH}
\mathrm{O}$Denatured Alcohol: To prevent the misuse or wrong use of ethanol produced for industrial use, it is made unfit for use that is for drinking by adding poisonous substances like methanol to it and also dyes to give it a fake colour. Atoms of other elements like hydrogen, oxygen, nitrogen, chlorine also show sharing of valence electrons. The
first fullerene identified was C-60 with 60 carbon atoms arranged like the geodesic dome designed by US architect, Buckminster Fuller, hence these are also known as Buckminster Fuller, hence the second as a supplication of the second as a supplicatio
with long chain. This property is due to The small size of the carbon atom. If the name of the current is deleted and the suffix is added. Ethanol alcohol reacts with sodium(Na) metal vigorously to form sodium ethoxide and evolves H2 gas. The earth's crust has only 0.02% to be given as a suffix, the last letter 'e' in the name of the current is deleted and the suffix is added. Ethanol alcohol reacts with sodium(Na) metal vigorously to form sodium ethoxide and evolves H2 gas. The earth's crust has only 0.02% to be given as a suffix, the last letter 'e' in the name of the suffix is added.
carbon in the form of minerals (like carbonates Class 10 Science Chapter 4 Bonding in Carbon: The Covalent
bond, Electron dot structure, Physical properties of organic compounds, Allotropes of Carbon. Catalyst: Substances that cause a reaction to occur or proceeds to different rate without consuming in it are called a catalyst. Advantage of Detergents: The main advantage of detergent over soaps is that soaps cannot be used in hard water for washing
because hard water reacts with soap to form curdy white precipitate called Scum. 10. Electron Dot Structure of Hydrocarbons Isomerism: Compounds having the same molecular formula but different structural formula and properties are known as Isomerism. e.g., If a number of carbon atoms is three, then
the word root is a prop. Carboxylic acid with three carbons is propanoic acid. In the presence of alkaline KMnO4, it is oxidised to ethanoic acid. Double covalent bond: Double bond is formed by sharing of four electrons, i.e two pairs of electrons. General formula: RCOO-Na+ Detergent: Ammonium and sulphonate salts of long chain fatty acids are
called Detergent. Go Back to Class 10 Science Notes Chapter 4 carbon and its compounds class 10 formulates 6-8 marks in the final exam and is one of the carbon-containing cluster of 60 carbon atoms joined together to form spherical molecules. The soap
micelles helps in dissolving the dirt in water and wash our clothes. Functional Group: An atom or a group of atoms which make carbon compound reactive and decides its properties are called functional groups. Thus,
carbon overcomes this problem by sharing of its valence electrons with other carbon atoms or with atoms of other elements. It is difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to either gain or lose electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would be difficult for an atom of carbon to electrons as it would 
(CH3COOH) 290 391 Chloroform (CHCl3) 209 334 Ethanol (CH3CH2OH) 156 351 Methane (CH4) 90 111 Allotropy: The phenomenon in which the element exists in two or more different physical states with similar chemical properties are called Allotropy. The reaction is known as saponification. It bums with a blue flame in the
presence of O2 of air. This is saponification as soap is prepared by this. Formation of Hydrogen = 1 Number of valence electrons = 1 Formation of Hydrogen = 1 Number of valence electrons = 1 Formation of Hydrogen = 1 Number of valence electrons = 1 Formation of Carbon = 6 [2, 4] Number of valence electrons = 1 Formation of Hydrogen = 1 Number of valence electrons = 1 Formation of Carbon = 6 [2, 4] Number of valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Number of Valence electrons = 1 Formation of Hydrogen = 1 Formation of Hydr
CO2 Molecule Atomic number of Carbon = 6 [2, 4] Number of valence electrons = 6 Physical Properties of Organic Compounds Most of the organic compounds have low boiling
and melting point, due to the weak force of attraction (i.e., the inter-molecular force of attraction) between these molecules. Carbon, hence, overcomes this difficulty by sharing it's four valence electrons with other atoms of carbon or with atoms of other elements. Revising notes in exam days is on of the best tips recommended by teachers during
exam days. CBSE quick revision note for Class-10 Science, Chemistry, Maths, Biology and other subject are very helpful to revise the whole syllabus during exam days. Saturated Hydrocarbon (Alkanes): General formula is CnH2n+2. Characteristics of a homologous series and the revision note for Class-10 Science, Chemistry, Maths, Biology and other subject are very helpful to revise the whole syllabus during exam days.
formula. For example, compounds of carbon with hydrogen are called hydrocarbons. Presence of a functional group is indicated by prefix or suffix as given in table 2, and table 3. 

Carbon and its Compounds MindmapsBelow, we have provided Class 10 Science Ouestion Banks that cover every typology question with detailed explanations from
various resources in one place. > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds Kendriya Vidyalaya Question Bank > Class 10 Carbon and Its Compounds Kendriya Vidyalaya Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its Compounds CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Question Bank > Class 10 Carbon and Its CBSE Quest
the number of electron pairs shared, covalent bond is of three types: Single Covalent bond is formed because of sharing unsaturated fatty acids are good for health. For instance, on ignition of wood or charcoal, the volatile substances
present in them vapourise and burn with a flame. Sweet smelling ester is formed. • Tetravalency: Carbon has 4 valence electrons. For example; Ni, Pt, Pd, etc. Use of detergents overcome this problem as the detergent molecule prevents the formation of insoluble product and thus clothes get cleaned. Carbon and its Compounds class 10 Notes CBSE
Revision notes (PDF Download) Free CBSE Revision notes for Class 10 Science PDF Revision notes and Key Points Class 10 Science PDF Revision notes for CBSE Revision notes and Key Points Class 10 Science PDF Revision notes for CBSE Revision notes and Key Points Class 10 Science PDF Revision notes for CBSE Revision note
notes for CBSE board exams CBSE Class 10 Revision Notes, sample paper for class 10 Mathematics, Science English Communicative; do check myCBSEguide app or website. For example, a compound with 4 carbon
atoms would have the name butane. If functional group is present, the name of the functional group is to be given as a suffix, and if suffix of the functional group is to be given as a suffix, and if suffix of the functional group is to be given as a suffix, and if suffix of the functional group is present, the name of the functional group is to be given as a suffix, and if suffix of the functional group is to be given as a suffix of the functional group is to be given as a suffix of the functional group is to be given as a suffix of the functional group is present, the name of the functional group is given based on the functional gro
the final 'e' with the appropriate suffix. If the octets are incomplete, and more electrons remain to be shared, move one electron per bond per atom to make another bond. Hence, detergents are used to overcome this problem, as detergents are sodium salts of sulfonic acids or ammonium salts with chlorides or bromide ions. (iv) Substitution Reaction
Important Carbon Compounds: Ethanol and Ethanoic acid Physical Properties of Ethanoic acid Physical 
sample papers with solution, test papers for chapter-wise practice, NCERT solutions, NCERT exemplar solutions, quick revision notes for ready reference, CBSE guess papers and CBSE important question papers. Members of given homologous series have the same functional group. Homologous series of organic compounds in which every
succeeding member differs from the previous one by -CH2 group or 14 a.m.u. Note: As the molecular mass increases in a series, : so physical properties of the compounds show a variation, but chemical properties which are determined solely by a functional group, remains same within a series. Therefore, carbon forms four covalent bonds, i.e.,
carbon is tetravalent. • Any two homologues differ by - CH2 group and difference in molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot structures) are diagrams that show the bonding between atoms of a molecular mass is 14µ. Example: Electron Dot structures (\electron Dot
emulsion is thus formed by soap molecule. Diamond is a bad conductor of electricity but graphite is a very good conductor of electricity. Carbon always forms a covalent bond: Atomic no of carbon is 6. Name the Compounds By Following Order Prefix + Word Root + Suffix Chemical Properties of Carbon Compounds: The important chemical properties
are as follows: 1. When the surface of the cloth is mechanically scrubbed or beaten on a stone or with a wooden paddle or agitated in a washing machine, the loosened oily particle is removed from the dirty surface and the cloth is cleaned. The solution now contains small globules of oil surrounded by detergent molecules. 16. Install NowCBSE class
10 Science Chapter 4 Carbon and its Compounds notes in PDF are available for free download in myCBSEguide mobile app. Combustion: The atoms of other elements like hydrogen, oxygen and nitrogen, chlorine also form bonds by sharing of
electrons. \rightarrow They are generally poor conductor of electricity as electrons are shared between atoms and no charged particles are formed. It is sour in taste. (iii) Their efficiency decreases in hard water (iii) Their efficiency decreases in hard water (iii) Their efficiency is unaffected in hard water. It is sour in taste.
most metallic or least electronegative atom in the center. Double Covalent Bond: When two pairs of electrons are shared between two atoms in a molecule. ADDITION REACTIONHydrogen is added to unsaturated hydrocarbon in presence of palladium or nickel as catalyst. The gain of 4 electrons (to form an octet, i.e., 8 electrons in C4- anion) is
difficult because then a nucleus with 6 protons will have to hold extra four electrons. For example; Methane (C44), Ethane (C46) etc. Alcohol (-OH) — ol 6. (i) These are ammonium and sulphonate salts of long chain fatty acids. These are heteroatom. Melting point and boiling point increases with increasing molecular mass. How to name carbon
compounds: NOMENCLATURE1. CH4(g) + 2O2(g) - CO2(g) + 2H2O(l) + Heat and light In presence of limited supply of air, saturated hydrocarbon forms a sooty flame. Like carbon, silicon called silanes. For example; O2, CO2 etc. The food we eat, the clothes we wear, everything
has some element of carbon in it. An unsaturated hydrocarbon with a triple bond between carbon atoms is named as Ethyne. Four carbon atoms is named as Ethyne. Four carbon in it. An unsaturated hydrocarbon with a triple bond between carbon atoms (4-C) But + ane 5. 5-8% solution of acetic acid in water is called Vinegar. Ethanol and Ethanoic Acid properties are discussed in-depth in this chapter. 

Worksheet 37 

Worksheet 38Below, we have
provided the links to downloadable PDFs of Experiential Learning Activity for class 10 Science to help students implement their acquired knowledge in the real world. > Carbon and its Compounds Experiential Activities Below, we have provided Class 10 Science Important Questions that cover all the important questions in Carbon and its
Compounds. • H - H single bond between hydrogen atoms (H2) • O = O double bond between one oxygen atoms • Water molecule has single covalent bond between one oxygen atoms • Water molecule has single covalent bond between nitrogen atoms • Water molecule has single covalent bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms (H2) • O = O double bond between oxygen atoms
When soap or detergent is dissolved in water, the molecules gather together as clusters, called micelles. (i) Saturated Hydrocarbons ice like solid. Hydrocarbons: Compounds of carbon and hydrogen are known as hydrocarbons. ◆
Some Important Carbon Compounds: Ethanol (CH3CH2—OH): Commonly known as Ethyl Alcohol. Class 10 Science notes on Chapter 4 Carbon and its Compounds Class 10 notes ScienceDownload CBSE class 10th revision notes for Chapter 4 Carbon and its
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