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How to find group number in periodic table

It's human nature to organize things. Cooks painstakingly organize their spices into various groupings, whether alphabetically or according to how often they're used. Kids dump out their piggy banks and sort their riches into piles of pennies, nickels, dimes and quarters. Even the items in a grocery store are grouped a certain way. Head down the international aisle, and you'll find packages of Chinese egg noodles sitting next to boxes of taco shells.Chemists, as it turns out, are organizational junkies, too. They look for similar physical and chemical properties among the elements, the basic forms of matter, and then try to fit them into similar groups.Scientists began attempting to organize the elements in the late 1800s when they knew of about 60. Their efforts, however, were premature since they were missing a key piece of information: the structure of the atom. While initial efforts failed, one attempt by a Russian chemist named Dmitry Mendeleev showed much promise. Although Mendeleev wasn't 100 percent correct, his approach laid the groundwork for what is now the modern periodic table of the elements.Today, the periodic table organizes 112 named elements and acknowledges several more unnamed ones. It has become one of the most useful tools in chemistry, not only for students, but for working chemists as well. It classifies the elements according to their atomic number (more on that soon), tells us about the nuclear composition of any given element, describes how electrons are arranged around a given element and allows us to predict how one element will react with another.So, exactly what is this feat of organization? Keep reading as we examine the history, organization and uses of this most handy chemical tool. Keep up with the latest daily buzz with the BuzzFeed Daily newsletter! Just the thought of looking at the periodic table makes us cringe because of everything that has to be remembered. But thinking about how it has helped us know so much information by just looking at it, gives us some sense of relief. So, thinking about it that way, periodic tables aren't a confusing problem—they're a helpful problem-solver.What if we make things interesting and give the periodic table a fashion-forward twist? Sounds crazy, right?Well, we recently came across this really amazing periodic table by Sophie and Trey that every style enthusiast would love. The best part about this table is that it has everything you need to rid you from your everyday dressing woes, from color tones belonging from every category like jewel, neutral, bright and soft, to fashion articles to help you decide what goes well with what and how to accessorize your complete look. It's more like your virtual closet that answers the most important question that you don't normally get an easy answer for: "What should I wear today?" the help of this table, we decided to put you out of your sartorial misery and help you with deciding how you can dress yourself on a daily basis while keeping yourself connected with the trends. Well, they say you aren't completely dressed until you wear the right accessories, so why don't we see how these everyday outfits can be matched with the right accessories (by the way, they don't actually say that, but you need accessories to complete your look along with smile, so why not).Yes, there's an easy equation for this. If you have the essentials sorted, the rest is a piece of cake. You can pair your favorite top with your least favorite bottom, go with an all-black everything look and can also accessorize a rather casual outfit, making it look fancy. So, let's give you the equation:Basic + Dressy + Fixer + Accessories = Your Wardrobe for Every day Now that you have a clear equation, let's work on unlocking this achievement of dressing up flawlessly every day without much effort.1. Begin with the basicsYour basic section of the wardrobe is usually based on neutral colors and staples pieces. If you take a look at the table, a number of ideas to adorn your basic clothing needs will start popping in your head. A basic top in ivory, pencil skirt in charcoal grey accessorized with an ink blue pair of pointed toe heels and a satchel or messenger will make up for an effortless work look. If you are looking for a more exuberant look for your everyday needs, build your basic wardrobe with a mix of bright and neutral or bright and soft tones.What you'll need Basic tops and blouses Sweat shirts Tees Pencil skirts Maxi skirts Culottes Tailored pants Cropped pants A pair of dark washed and light washed jeans Trend alertCasual and minimal is actually the new black for this season. However, the rule has a stylish shift to it: the simpler the outfit, the dressier the accessories. However, for an everyday look, keep it subtle, without making anything look out of proportion. The table has all the right accessories for you to pick from. Keeping up with trends, the basic white button down is definitely a thing because the runways had them all, from shirts, dresses, to coats and jacket. The classic button down surely has elevated its way to the top.The pantone colors of this year: serenity and rose quartz can easily be integrated in your basic wardrobe because of their soothing yet versatile appeal. The colors can be worn both formally and casually. 2. Deal with dressyThis is the category where you can be as creative as you want. You can get your hands on colors, prints, textures, patterns, shimmer or just about anything that can easily work as fancy or vivid. For your rather dressier looks, go with jewel and bright tone. If you really need to turn the colors down, then go for pairing them with soft tones. Tangerine, blood orange, cranberry and emerald green are colors that have a quintessential appeal to them and are just the right hues for night and evening wear.What you'll need Dresses Skirts Bold prints Sequins Fringes Maxi Dresses LBDs LWDs Skirts Gowns Slip dresses Trend alertThere still is a love affair with the 70s going on in the fashion world. Designers for the season were particularly interested in incorporating the 90s and 70s with some modern alteration in their creations. Voluminous ruffles, Victorian and English floral prints, puffy sleeves, cropped details, fringed sequins, and the return of the slip dress all indicates that there is some serious 70s renaissance is happening.Make sure you have some romantic floral prints and dresses with bold stripes and solid color play hanging in your dressier section of the closet. Oh, and you may have not noticed but silk PJ's with boisterous prints are now worn formally. So, if you really are going to go with the trends, choose the accessories and tones from the periodic table and invest in some glamorous silk PJ's. A Spanish spin to prints and cuts was witnessed as a key trend on the runways; you can wear the trend through Spanish-looking prints in your dresses.For all the good reasons, off-the-shoulder outfits managed to stay in vogue for the season. So, if you catch your eyes on a ruffled off-the-shoulder top, get that! 3. Fix it up with the fixersThe fixers are more like completers—the cherry on top. Don't you always need a layer, a spring jacket, a cardigan or a belt to finish off the outfit details? Well, that's what this little section of your wardrobe will be based on. You would not need much to add to this part but make sure you choose the right colors so that one fixer can work well with a number of outfits.What you'll need A light spring jacket Simple pretty cardigan Chunky belts Scarves Trend AlertYou will be surprised, but jackets and cardigans are still going strong in spring. In fact, they will be gracing the streets this year, more than ever. Why? Because that's what we're seeing in fashion shows and on runways. Jackets, sweatshirts, oversized hoodies will be all you need—yes, you can leave your pants at home. Pair them with knee-high boots and stockings to give your legs some cover. As for belts, they are the ultimate fixer and your friend in sartorial needs.4. Accessorize for perfectionYour style equation is never complete without the right shoes, jewelry and bags. So, carefully plan this portion according to the table. You have a number of neutral and bright colors to pick your shoes and bags from. Take a look at the accessories in the periodic table, and you'll know what you need for your day-to-day outfits.What you'll need Bags Clutches Heels Chunky sandals Flats Watch Statement earrings and necklaces Delicate ear studs and necklaces Trend AlertThe trends in accessories are more or less the same; shoes get sequined embellishment stone details. However, the more chunky they are, the trendier. Thankfully, we're still not being too hard on ourselves and going for more flats than heels. As for bags, invest in a chic half-moon bag and some bold colored bags with funky details. The jewelry has a standout element, but they are not heavy on the wearer, so you are free to wear those statement earrings and on a regular day, nothing's better than a statement watch and delicate jewelry.Stay stylish diva, and get the table printed and pated on your wardrobe to minimize the struggle.Colors this season transcend cultural and gender norms. Vivid brights give way to excitement and optimism, though quiet stability prevails in this season's paletteFeatured photo credit: Sophie and Trey via sophieandrey.com The horizontal rows on the periodic table of the elements are called periods. Every element in a period has the same number of atomic orbitals. For instance, hydrogen and helium are in the first period, so they both have electrons in one orbital. The columns on the table divide the elements into groups with the same number of electrons in their outer shells. These electrons, called valence electrons, cause them to share chemical properties. For example, noble gases have full valence shells, so they are unreactive, and halogens, such as fluorine and chlorine, react vigorously with alkali metals, such as sodium and potassium. Hydrogen is the element that is atomic number 1 on the periodic table. The element number or atomic number is the number of protons present in the atom. Each hydrogen atom has one proton, which means it has a +1 effective nuclear charge. At room temperature and pressure, hydrogen is a colorless, odorless gas. While ordinarily classified as a nonmetal, the solid form of hydrogen acts like other alkali metals in the same column of the periodic table. Hydrogen metal forms under intense pressure, so it is not seen on Earth, but it does exist elsewhere in the solar system. The pure element bonds to itself to form diatomic hydrogen gas. This is the lightest gas, although it is not significantly lighter than helium gas, which exists as a monatomic element. Element atomic number 1 is the most abundant element in the universe. In terms of a sheer number of atoms, about 90% of atoms in the universe are hydrogen. Because the element is so light, this translates into around 74% of the universe by mass. Hydrogen is extremely flammable, but it doesn't burn without the presence of oxygen. If you were to place a lit match into a container of pure hydrogen, the match would simply go out, not cause an explosion. Now, if it was a mixture of hydrogen and air, the gas would ignite! Many elements can exhibit a variety of oxidation states. While atomic number 1 usually displays a +1 oxidation state, it can also pick up a second electron and exhibit a -1 oxidation state. Because two electrons fill the s subshell, this is a stable configuration. There are three isotopes that all have atomic number 1. While an atom of each isotope has 1 proton, they have different numbers of neutrons. The three isotopes are proton, deuterium, and tritium. Protium is the most common form of hydrogen in the universe and in our bodies. Each protium atom has one proton and no neutrons. Ordinarily, this form of element number 1 has one electron per atom, but it readily loses it to form the H+ ion. When people talk about "hydrogen", this is the isotope of the element usually being discussed. Deuterium is a naturally occurring isotope of element atomic number 1 that has one proton and also one neutron. Since the number of protons and neutrons is the same, you might think this would be the most abundant form of the element, but it's relatively rare. Only around 1 in 6400 hydrogen atoms on Earth are deuterium. Although it's a heavier isotope of the element, deuterium is not radioactive. Tritium also occurs naturally, most often as a decay product from heavier elements. The isotope of atomic number 1 is also made in nuclear reactors. Each tritium atom has 1 proton and 2 neutrons, which is not stable, so this form of hydrogen is radioactive. Tritium has a half-life of 12.32 years. Helium is the element that is atomic number 2 on the periodic table. Each helium atom has 2 protons in its atomic nucleus. The atomic weight of the element is 4.0026. Helium does not readily form compounds, so it is known in its pure form as a gas. Element Name: Helium Element Symbol: He Atomic Number: 2 Atomic Weight: 4.002 Classification: Noble Gas State of Matter: Gas Named For: Helios, the Greek Titan of the Sun Discovered By: Pierre Janssen, Norman Lockyer (1868) The element is named for the Greek god of the sun, Helios, because it was initially observed in a previously unidentified yellow spectral line during the 1868 solar eclipse. Two scientists observed the spectral line during this eclipse: Jules Janssen (France) and Norman Lockyer (Britain). The astronomers share credit for the element discovery. Direct observation of the element did not occur until 1895, when Swedish chemists Per Teodor Cleve and Nils Abraham Langlet identified helium emanations from cleveite, a type of uranium ore. A typical helium atom contains 2 protons, 2 neutrons, and 2 electrons. However, atomic number 2 can exist without any electrons, forming what is called an alpha particle. An alpha particle has an electrical charge of 2+ and is emitted during alpha decay. The isotope containing 2 protons and 2 neutrons is called helium-4. There are nine isotopes of helium, but only helium-3 and helium-4 are stable. In the atmosphere, there is one atom of helium-3 for every million helium-4 atoms. Unlike most elements, the isotopic composition of helium greatly depends on its source. So, the average atomic weight may not really apply to a given sample. Most of the helium-3 found today was present at the time of the Earth's formation. At ordinary temperature and pressure, helium is an extremely light, colorless gas. Helium is one of the noble gases or inert gases, which means it has a complete electron valence shell so it's not reactive. Unlike gas of atomic number 1 (hydrogen), helium gas exists as monatomic particles. The two gases have comparable mass (H2 and He). Single helium atoms are so small they pass between many other molecules. This is why a filled helium balloon deflates over time -- the helium escapes through tiny pores in the material. Atomic number 2 is the second most abundant element in the universe, after hydrogen. However, the element is rare on Earth (5.2 ppm by volume in the atmosphere) because nonreactive helium is light enough that it can escape Earth's gravity and be lost to space. Some types of natural gas, such as that from Texas and Kansas, contain helium. The primary source of the element on Earth is from liquefaction from natural gas. The largest supplier of the gas is the United States. The source of helium is a non-renewable resource, so there may come a time when we run out of a practical source for this element. Atomic number 2 is used for party balloons, but it's primary use is in the cryogenic industry for cooling superconducting magnets. The principal commercial use of helium is for MRI scanners. The element is also used as a purge gas, to grow silicon wafers and other crystals, and as a protective gas for welding. Helium is used for research into superconductivity and the behavior of matter at a temperature approaching absolute zero. One distinctive property of atomic number 2 is that this element cannot be frozen into a solid form unless it is pressurized. Helium remains liquid down to absolute zero under normal pressure, forming a solid at temperatures between 1 K and 1.5 K and 2.5 MPa pressure. Solid helium has been observed to possess a crystalline structure. Hammond, C. R. (2004). The Elements, in Handbook of Chemistry and Physics (81st ed.). CRC press. 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