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Fractional distillation of crude oil pdf

Raw oil (petroleum) is made up of hydrocarbons mostly (alkanes and aromatics), and is a huge mixture of compounds composed between 5 and 40 carbon atoms.⁽¹¹⁾ Oil components are incredibly useful as fuels and lubricants, but not when mixed together. Fractional distillation is used in oil refineries (Figure 5.41) to separate the complex mixture into fractions containing similar boiling points and therefore similar molecular weights and properties. Gasoline, diesel fuel, kerosene and jet fuel are some of the different fractions produced by an oil refinery. Figure 5.41: San Francisco Raffineria in Rodeo, California. Cyclpentadiene is used in many chemical reactions, including Diels-Alder reactions and polymerizations. The reagent is so reactive, however, that it undergoes a reaction of Diels-Alder with it in the reagent bottle to form the dicyclopentadiene (Figure 5.42a). Therefore, chemical companies do not sell cyclopentadiene, and chemicals are instead required to distill commercial dicyclopentadiene (Figure 5.42b) to reverse the dimerization reaction and get cyclopentadiene (Figure 5.42c). Figure 5.42: a) Dimerization reaction of cyclopentadiene, b) Dicyclopentadiene reagent, c) Reaction of Diels-Alder Retro. At temperatures above 150°C the reaction is reversed. The dimer returns to the monomer through a reaction retro Diels-Alder (guided by favorable change in entropy, Figure 5.42c). Distillation can be used to remove monomer as a form. The two components (dimers and monomers) have remarkably different boiling points, the temperature required for the reverse reaction is too similar to the boiling point of the dicyclopentadiene which its steam pressure cannot be ignored. For this reason, fractional distillation is necessary for this process. 6% of crude oil contains hydrocarbons with more than 40 carbon atoms, a fraction that is eventually used for asphalt. Contributor Lisa Nichols (Butte Community College). Organic Chemistry Laboratory Techniques is licensed under a Creative Commons Attribution-Non Commercial-NoDerivatives 4.0 International License. Full text is available online. Most of our fuels and plastics are derived from oil. Raw oil can be separated into different fractions using fractional distillation. The distillation is used to separate liquid mixtures using differences in the boiling points of the different components. The technique is widely used in industry, for example in the manufacture and purification of nitrogen, oxygen and rare gases. However, one of its best known uses is the refinement of crude oil in its main fractions, including naphtha, cherosin and gas oil. This is described below. This is the first phase of oil conversion in compounds that are then used to produce everything from plastics to drugs. Figure 1 Carling's steam cracking plant (near Metz in eastern France) is dominated by distillation towers. They produce ethene and propene that are converted into poly(ethene) and poly(propene). Please grant Total. Crude Oil Distillation Crude oil is a mixture of many hundreds of liquid hydrocarbons. Dissolved in it are many other hydrocarbons some of which are solid and some gases (the lower members of the alkan family, mainly methane and ethane, but often with some propane and butane). There may also be some hydrocarbon gases trapped above oil, such as in some of the North Sea oil fields. In the refineries the oil is distilled in liquid fractions with different boiling points intervals which are then further transformed. The crude oil is heated in an oven (ca 650 K) and the resulting mixture fed as steam in a fractional tower that can have a height of 25–100 m, handling volumes of more than 40 000 m³ per day. The column can contain steel axes 40–50 that adapt horizontally through the column and are designed to ensure that there is intimate mixing between the descending liquid, formed by condensation, and rising steam. In order to make this close contact, the trays have holes in them (<seed>) through which the steam flows into the liquid that collects on the trays (Figure 2). Figure 2 Trays in a fractional column. Alternatively, there can be valves that fit the tray, which lift when the steam pressure under the tray is greater than the pressure on the tray (Figure 7). These are considered more in fractionation compared to sieve trays without valves. Figure 3 trays valves. In the tower there is a temperature gradient; the floor is cooler than the bottom, when the rising steam reaches a tray containing liquid whose temperature is below the boiling point (bp) of the steam, it is partially condensed. Since some of the vapors condense to a liquid, the heat dissipated latent then heats more liquid, and the most volatile components in the liquid evaporate by combining the remaining steam and passing the tower, the less volatile liquid flows through the tray and down a pipe to the tray below. Figure 4 this flexi tray valve tray is a sheet of steel on which the lifting valves are mounted, are much more efficient than sieve trays. Courtesy of Koch-Glitsch, Inc. This process occurs continuously in each tray, less volatile steam components and more volatile evaporation. This results in each tray containing products with a relatively narrow boiling point (called "closed cut" of products.) This leads to molecular mass products (low bp) accumulating near the top of the tower and high molecular mass constituents (high bp) collecting near the bottom (Figure 5.) Top uses the length of the carbon tower (bp 20 BOTTOM Table 1 Main fractions of crude oil primary distillation. Figure 9 The fractional distillation of crude oil. The high boiling point residue from crude oil is then transferred to another column and distilled under vacuum; lowering the pressure reduces the boiling point and ensures distillation constituents at temperatures below the decomposition temperature. Lubricant oils and waxes are obtained from this process. The final residue of the process is bitumen. The details of the fraction are shown in Table 1. The constituents of crude oil vary from one oil field to another (for example, the percentage of naphtha from North Sea oil is considerably greater than Middle East oil) so that the data numbers are approximate. Last modified: 14 March 2016 fractional distillation of crude oil products. fractional distillation of crude oil steps. fractional distillation of crude oil pdf. fractional distillation of crude oil diagram. fractional distillation of crude oil ppt. fractional distillation of crude oil igcse. fractional distillation of crude oil experiment. fractional distillation of crude oil temperatures

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